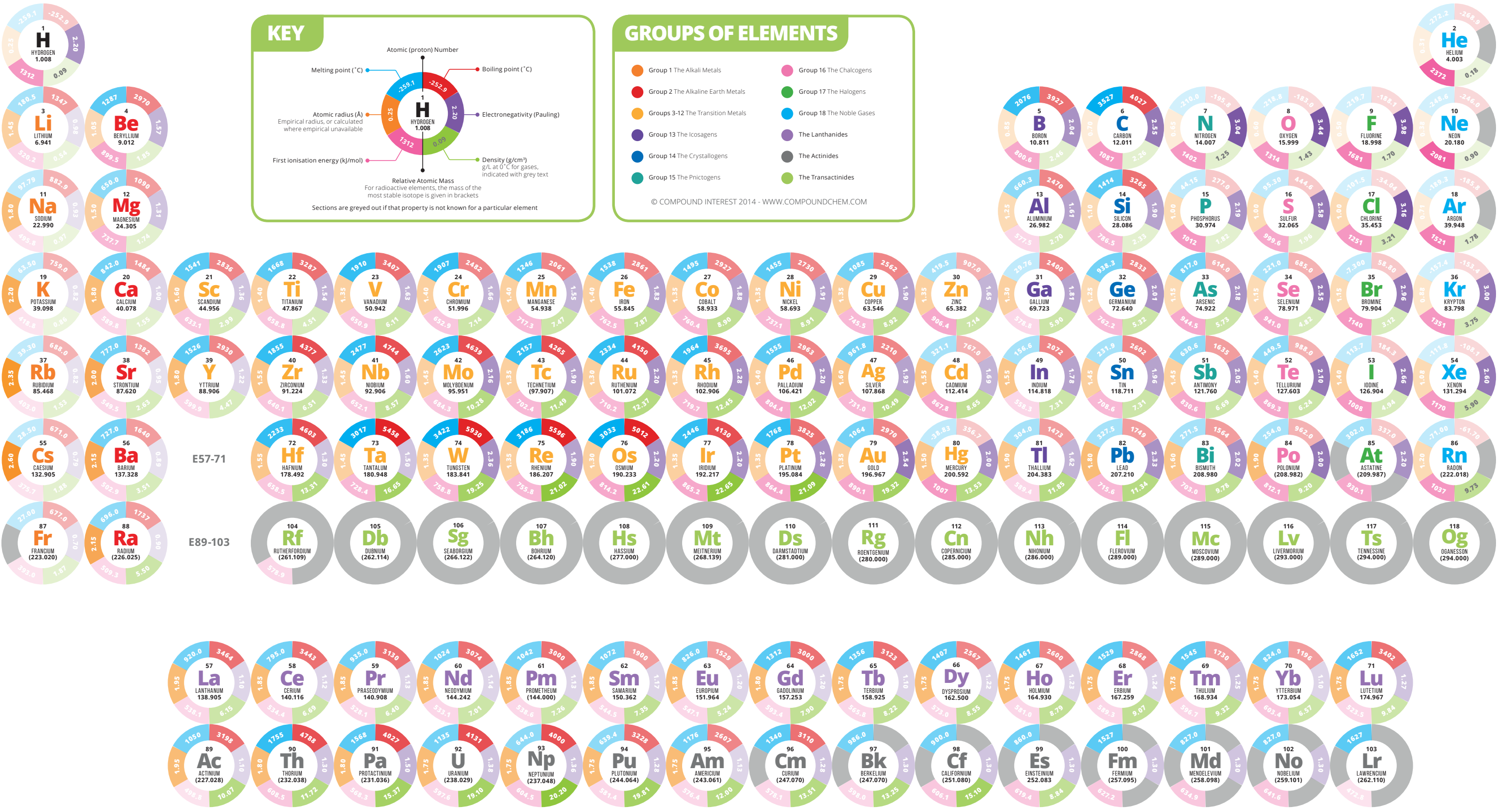


# THE PERIODIC TABLE OF THE ELEMENTS

1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18



### KEY

Atomic (proton) Number

Melting point (°C)

Boiling point (°C)

Atomic radius (Å)  
Empirical radius, or calculated where empirical unavailable

Electronegativity (Pauling)

First ionisation energy (kJ/mol)

Density (g/cm³)  
g/L at 0°C for gases, indicated with grey text

Relative Atomic Mass  
For radioactive elements, the mass of the most stable isotope is given in brackets

Sections are greyed out if that property is not known for a particular element

### GROUPS OF ELEMENTS

- Group 1 The Alkali Metals
- Group 2 The Alkaline Earth Metals
- Groups 3-12 The Transition Metals
- Group 13 The Icosagens
- Group 14 The Crystallogens
- Group 15 The Pnictogens
- Group 16 The Chalcogens
- Group 17 The Halogens
- Group 18 The Noble Gases
- The Lanthanides
- The Actinides
- The Transactinides

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NOTES: For elements with more than one allotrope, the properties of the most common allotrope are given. Different allotropes may have differing melting points, boiling points, and densities.