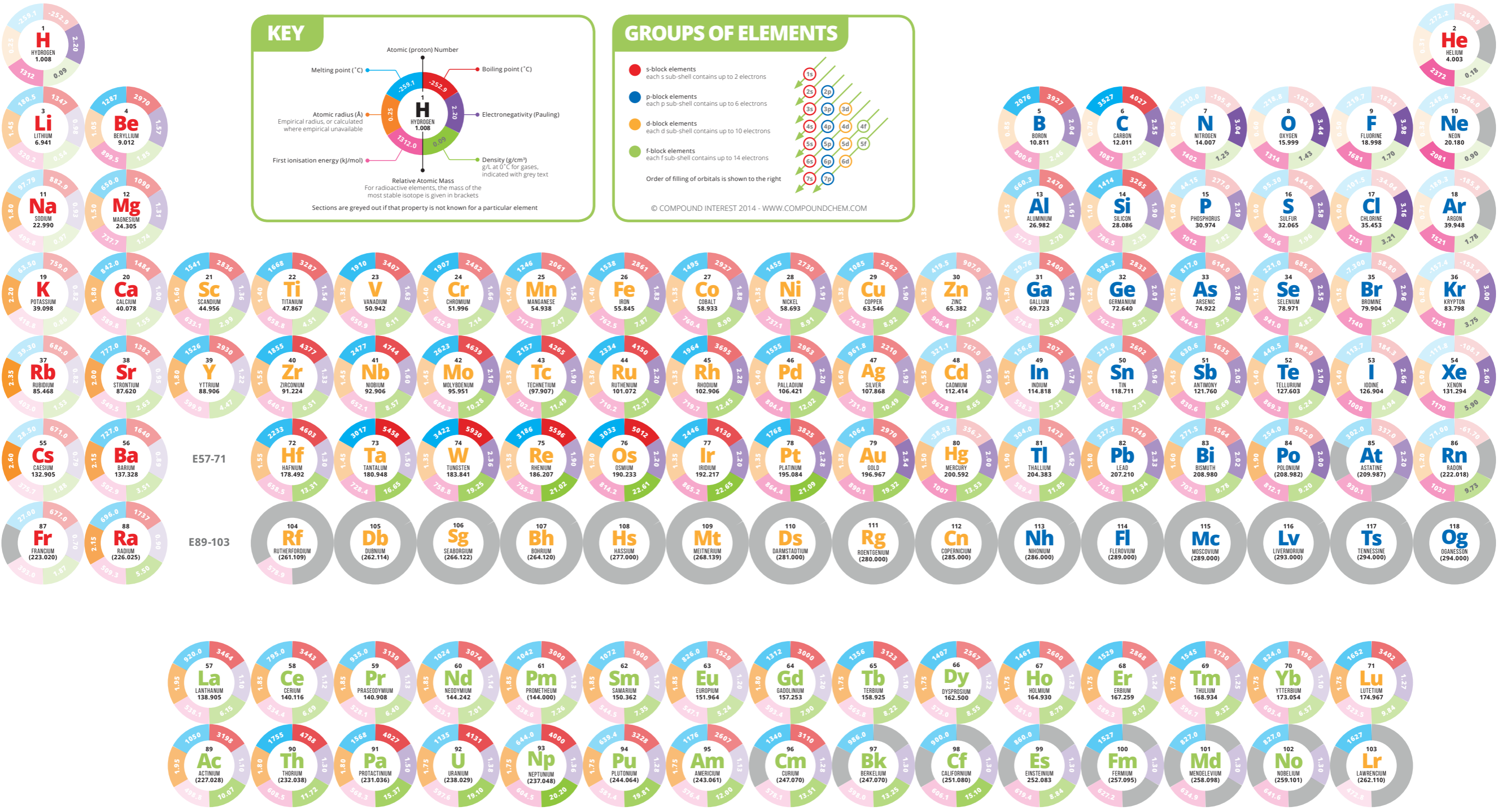


THE PERIODIC TABLE OF THE ELEMENTS

1 2 3 4 5 6 7 8 9 10 11 12 13 14 15 16 17 18



KEY

Atomic (proton) Number

Melting point (°C)

Boiling point (°C)

Electronegativity (Pauling)

Atomic radius (Å)
Empirical radius, or calculated where empirical unavailable

First ionisation energy (kJ/mol)

Density (g/cm³)
g/L at 0°C for gases, indicated with grey text

Relative Atomic Mass
For radioactive elements, the mass of the most stable isotope is given in brackets

Sections are greyed out if that property is not known for a particular element

GROUPS OF ELEMENTS

- s-block elements
each s sub-shell contains up to 2 electrons
- p-block elements
each p sub-shell contains up to 6 electrons
- d-block elements
each d sub-shell contains up to 10 electrons
- f-block elements
each f sub-shell contains up to 14 electrons

Order of filling of orbitals is shown to the right

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NOTES: For elements with more than one allotrope, the properties of the most common allotrope are given. Different allotropes may have differing melting points, boiling points, and densities.