

4
Be
Beryllium

12
Mg
Magnesium

20
Ca
Calcium

38
Sr
Strontium

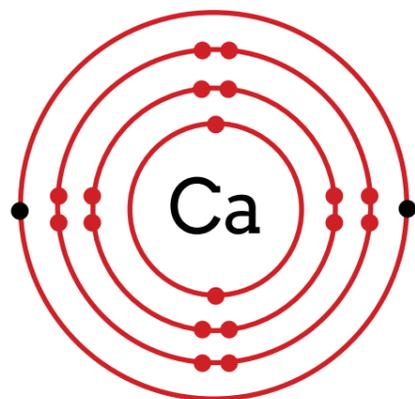
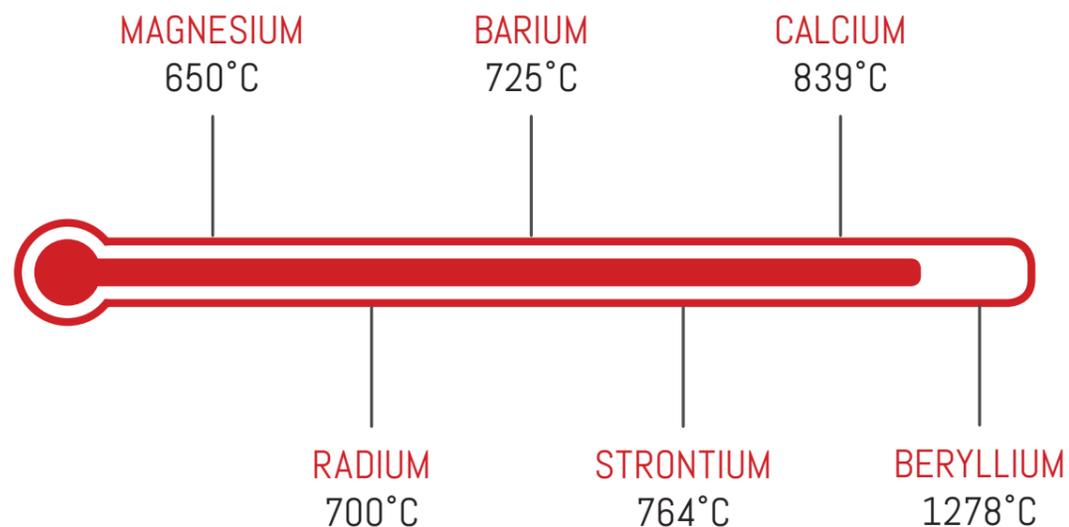
56
Ba
Barium

88
Ra
Radium

Group 2 - The Alkaline Earth Metals

THE GROUP 2 ELEMENTS ARE SHINY, SILVERY-WHITE, AND SOMEWHAT REACTIVE METALS, SOME OF WHICH OCCUR NATURALLY AS FREE ELEMENTS

MELTING POINTS



ALL OF THE GROUP 2 METALS HAVE TWO VALENCE ELECTRONS

THE REACTIVITY OF THE GROUP 2 METALS **INCREASES DOWN THE GROUP** AS THE OUTER ELECTRONS GET FURTHER FROM THE NUCLEUS & BECOME EASIER TO REMOVE
THEY ARE LESS REACTIVE THAN GROUP 1

THE ALKALINE EARTH METALS REACT WITH WATER TO FORM METAL HYDROXIDES...



EXCEPT FOR **Be** WHICH HAS A PROTECTIVE OXIDE LAYER PREVENTING REACTION

GROUP 2 METALS REACT WITH OXYGEN TO FORM METAL OXIDES

GROUP 2 METALS REACT WITH HALOGENS TO FORM METAL HALIDES

RADIUM



RADIUM IS A RADIOACTIVE ELEMENT WHICH USED TO BE USED TO MAKE GLOW IN THE DARK PAINT

USES OF THE ALKALINE EARTH METALS



BERYLLIUM
EMERALDS
TELESCOPE MIRRORS



MAGNESIUM
ALLOY WHEELS
FLARES



CALCIUM
BONES
BLACKBOARD CHALK



STRONTIUM
FIREWORKS
TREATING OSTEOPOROSIS



BARIUM
RAT POISON
GLASSMAKING