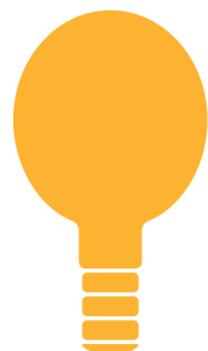


# The Transition Metals

A LARGE GROUP OF METALS IN THE CENTRE OF THE PERIODIC TABLE, THEY ARE LESS REACTIVE THAN THE GROUP 1 & 2 METALS, AND HAVE HIGH MELTING POINTS & DENSITIES

21 Sc Scandium	22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc
39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium
	72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury

**3422** °C  
**TUNGSTEN** HAS  
THE HIGHEST  
MELTING POINT  
OF ANY  
METALLIC  
ELEMENT



TRANSITION METALS FORM



COLOURED  
COMPOUNDS



THERE ARE  
THREE MAGNETIC METALS  
**IRON, COBALT  
AND NICKEL**

**MERCURY** IS THE ONLY  
LIQUID METAL AT  
ROOM TEMPERATURE **Hg**

UNLIKE GROUP 1 & 2 METALS  
TRANSITION METALS CAN FORM  
**CO-ORDINATION COMPLEXES**  
WITH OTHER ATOMS & MOLECULES

TRANSITION METALS  
HAVE  
VARIOUS OXIDATION STATES



MANGANESE HAS  
TEN POSSIBLE  
OXIDATION STATES

**COPPER, SILVER & GOLD**  
ARE KNOWN AS THE  
**COINAGE METALS**

