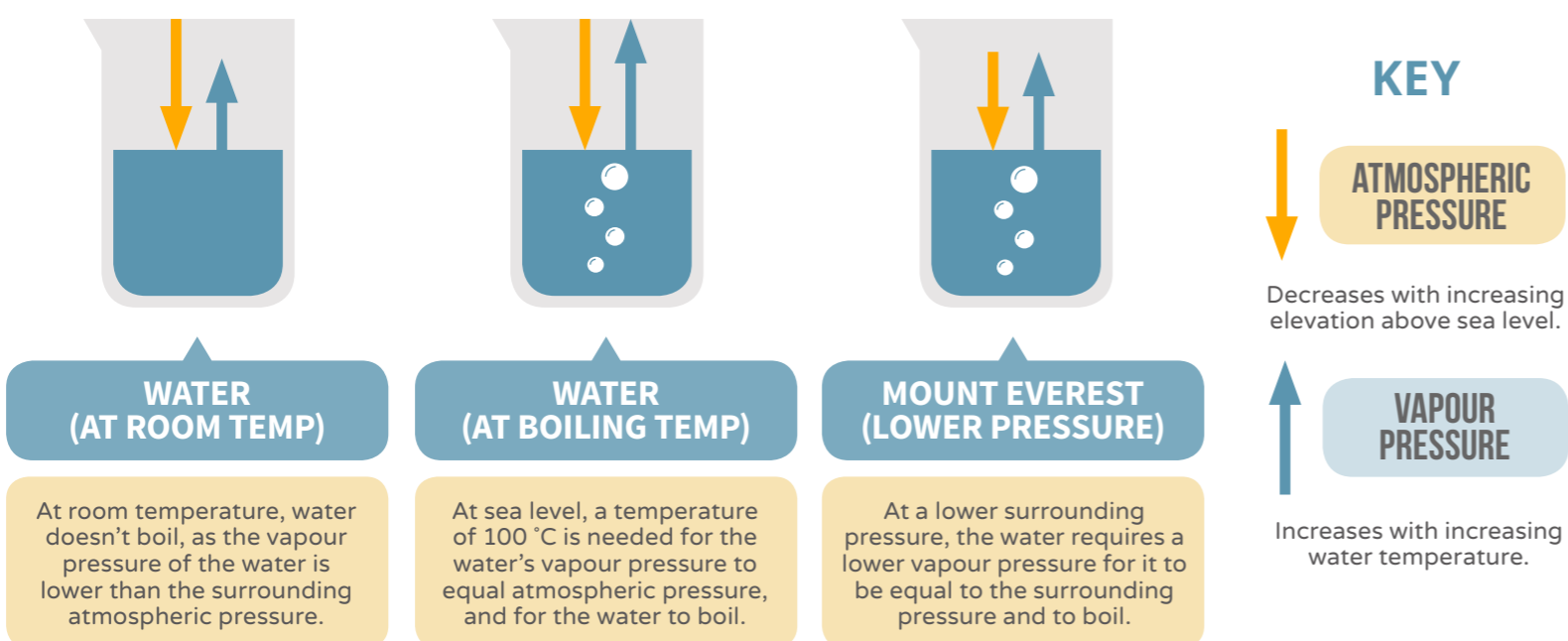
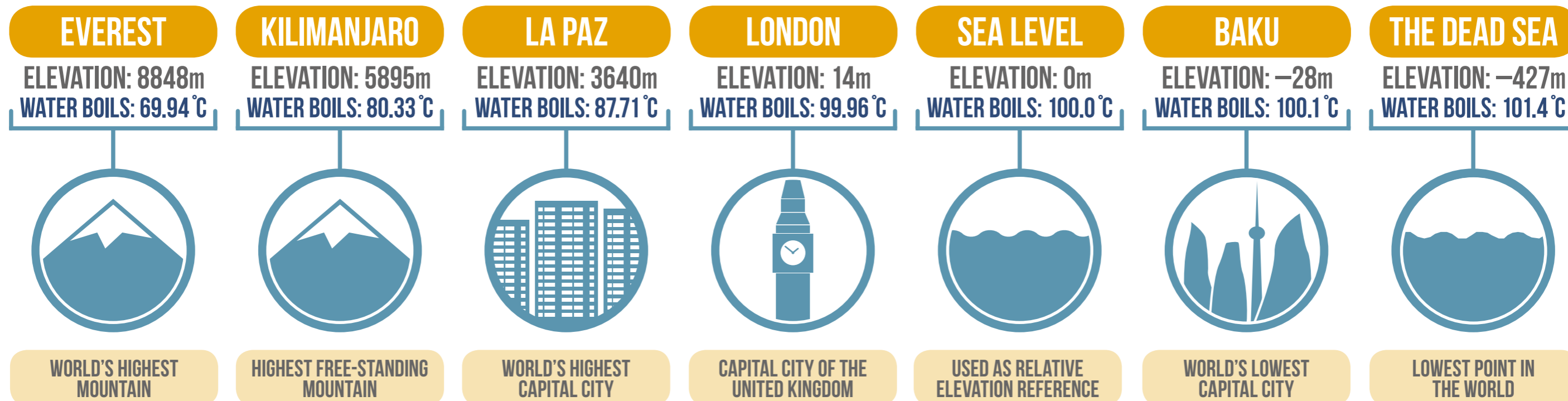


WHAT TEMPERATURE DOES WATER BOIL AT?

It might seem a pretty straightforward question – but actually, water's boiling point can differ at different elevations. This graphic takes a look at its boiling point in a several different locations, as well as looking at the reasons behind this variance.



WHY DOES WATER'S BOILING POINT VARY?

It's not so much the elevation that affects water's boiling point, as the decreased atmospheric pressure at higher elevations. A liquid will boil when its vapour pressure is equal to the atmospheric pressure; vapour pressure can be thought of as the tendency of molecules to escape the liquid's surface into the gas phase.

Vapour pressure increases with increased temperature, as more molecules have the kinetic energy required overcome attractions to other water molecules. At lower pressures, molecules escape more easily, as the vapour pressure required for them to do this is lower.

